

# HDO (and H<sub>2</sub>O/D<sub>2</sub>O) quenching rates

Theory & experiments

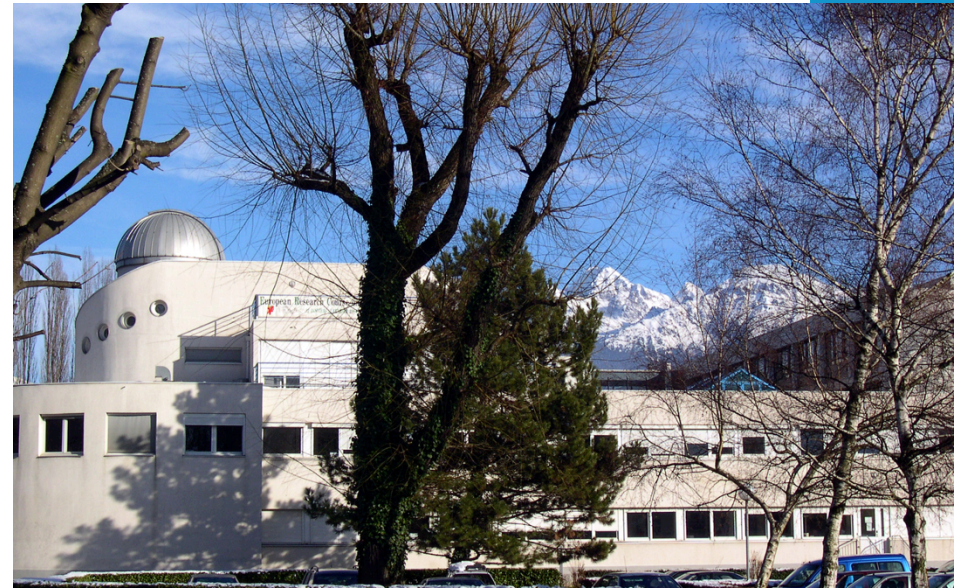
# A venture of IPAG and several institutes in France and worldwide

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Meudon, Le Havre, Dijon, Lille,  
Rennes, Bordeaux  
Nijmegen (NL), JPL-Pasadena (US),  
Durham (UK)

Also: ML Dubernet, F Daniel

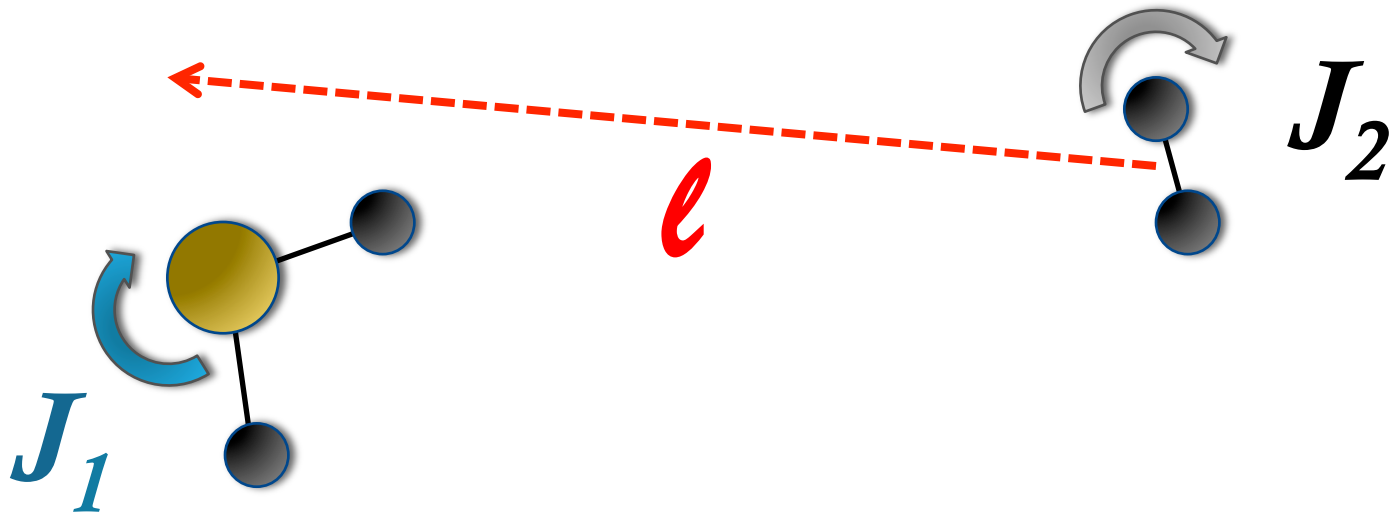
€ :  
EU FP6 'Molecular Universe',  
CHESS KP HSO CNES,  
ANR FORCOMS  
COST 'The Chemical Cosmos', PCMI-INSU-CNRS, LABEX  
OSUG@2020, NWO



# Collisional quenching/excitation :What is the physical process?

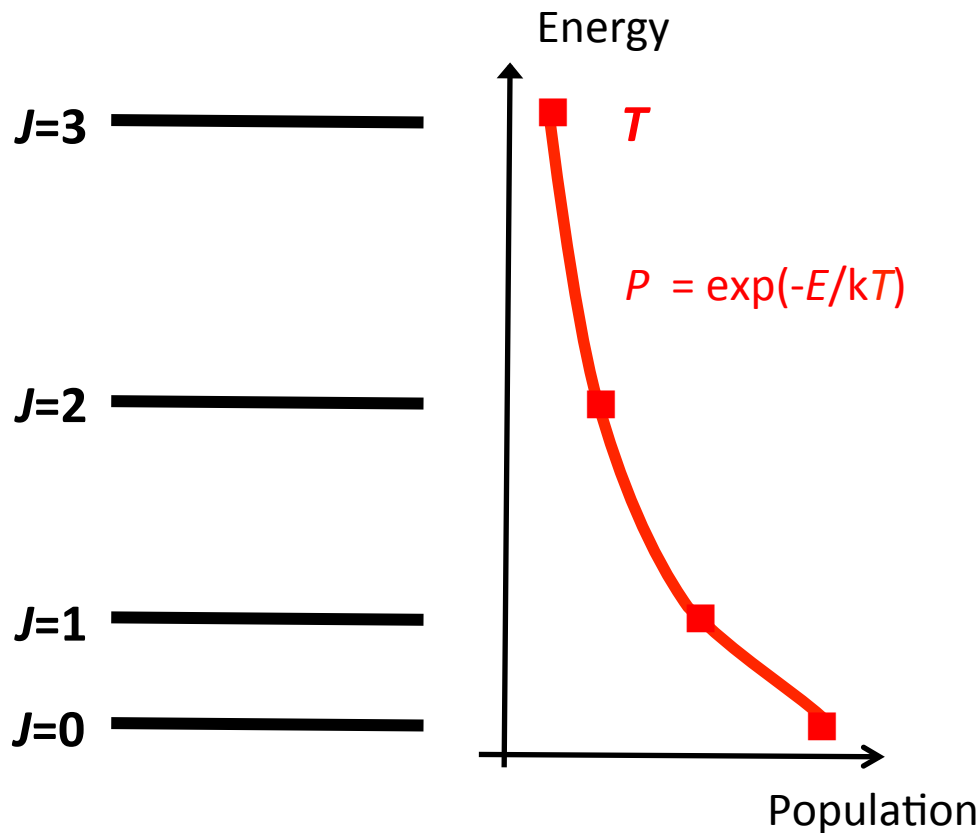
Rotational quenching (& fine structure, hyperfine, ro-vibrational quenching) :

***Transfer between projectile and target of kinetic energy /angular momentum (external and/or internal) + transient polarization***

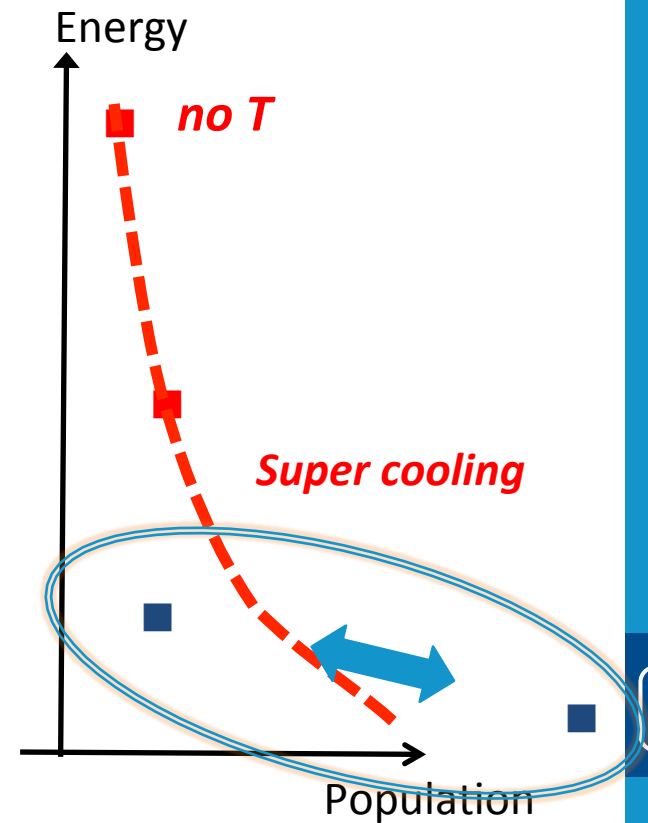


# Level populations

Levels and Temperature; Equilibrium with the photon gas or the buffer gas



Out of equilibrium : no  $T$ . Competition between photon interactions and collisions



# Dividing point: critical density

$$n^*(\text{H}_2) = \frac{\sum_{j < i} A_{ij}}{\sum_{j < i} k_{ij}} \quad (\text{cm}^{-3})$$

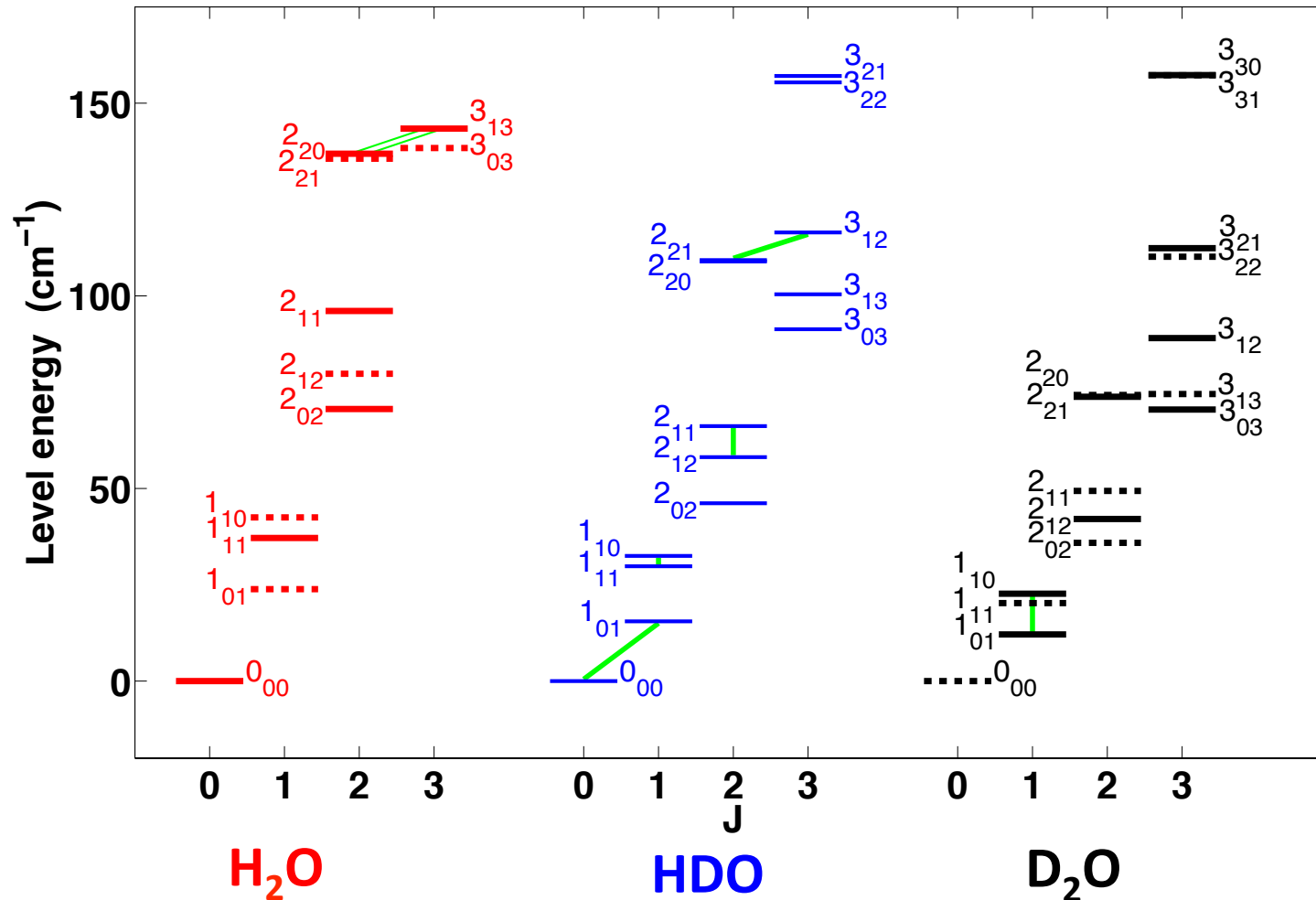
Spontaneous emission ( $\text{s}^{-1}$ )

Collision rates ( $\text{cm}^3 \text{s}^{-1}$ )

# Potential energy surface: *Water–molecular hydrogen* van der Waals interaction

- Full 9D PES, with
    - High precision 5D PES (rigid monomers), at CCSD(T)+ R12 corrections
    - Extrapolation for  $R > 15 \text{ \AA}$
    - Expansion around equilibrium points  $(x-x_{\text{eq}})^n$ ,  $n=2,3$
    - Average over  $v=0$  ground state wave function :  $V(\text{average})$ , 5D
    - Average over  $v>0$  wave functions : collision with excited  $\text{H}_2\text{O}$  (or  $\text{H}_2$ )
    - $\langle v=1 | V | v=0 \rangle$  matrix elements : vibrational quenching
- Valiron et al JCP, 2005 and 2009.***
- Monomer water : Kyrø Hamiltonian, ...

# Rotational levels H<sub>2</sub>O, HDO, D<sub>2</sub>O



# Using the full PES

Property	Isotope	Calculation
Cross section, rotational excitation	H <sub>2</sub> O – H <sub>2</sub>	Quantum/ semi-classical
Rates	H <sub>2</sub> O – H <sub>2</sub>	Quantum
		Classical/statistical
Cross sections,rates	<b>HDO – H<sub>2</sub></b>	<b>Quantum</b>
	D <sub>2</sub> O	Quantum
Differential Cross Section	H <sub>2</sub> O – H <sub>2</sub>	Quantum
	HDO-H <sub>2</sub> ; D <sub>2</sub> O – H <sub>2</sub>	Quantum
	H <sub>2</sub> O – D <sub>2</sub>	Quantum
Pressure Broadening	H <sub>2</sub> O – H <sub>2</sub>	Quantum
Vibrational excitation	H <sub>2</sub> O – H <sub>2</sub>	Classical, <b>quantum</b>



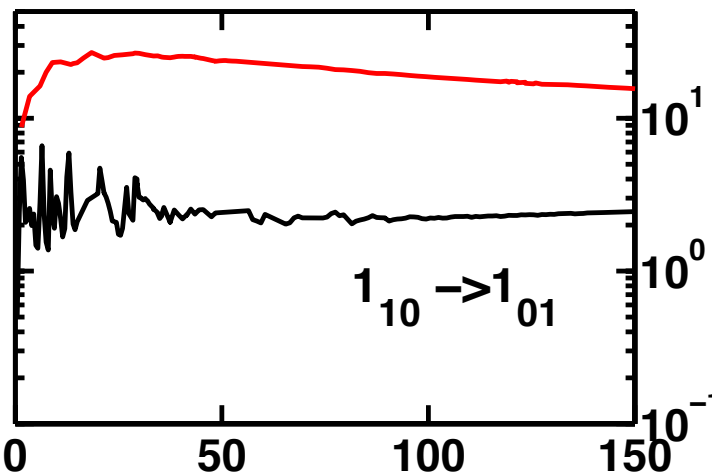
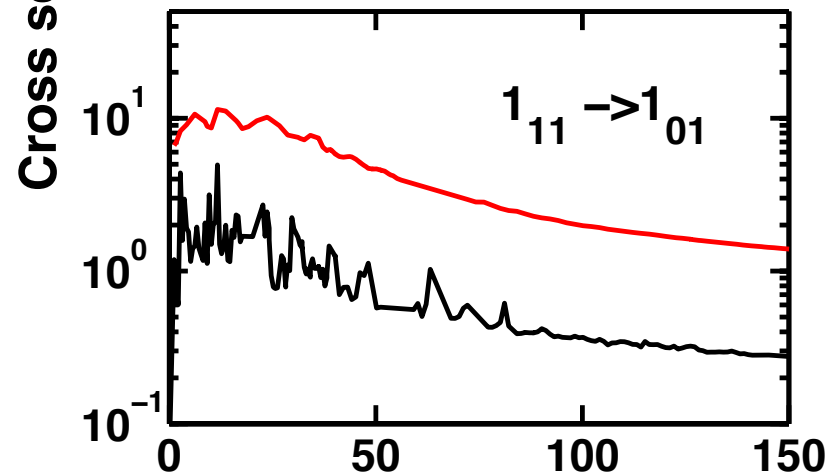
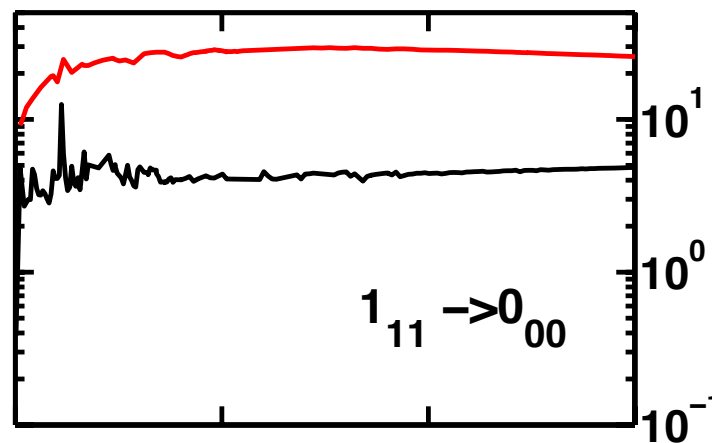
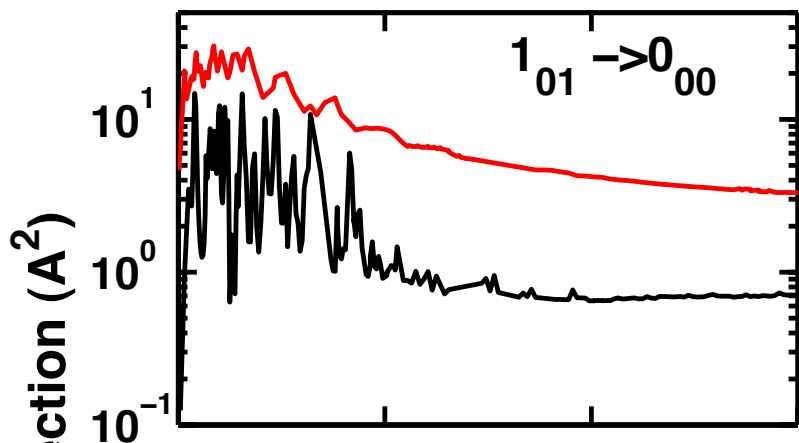
# Cross sections, HDO-H<sub>2</sub>

$$\sigma(E_{\text{collision}})$$

Ortho H<sub>2</sub>  $j=1$



Para H<sub>2</sub>  $j=0,2$



Collision energy (cm<sup>-1</sup>)

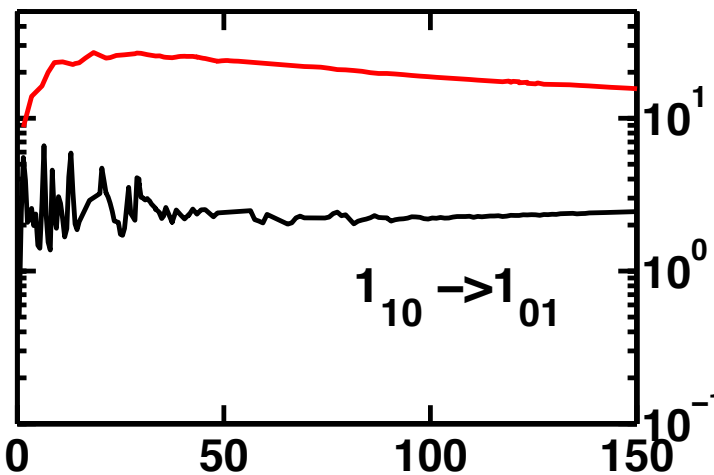
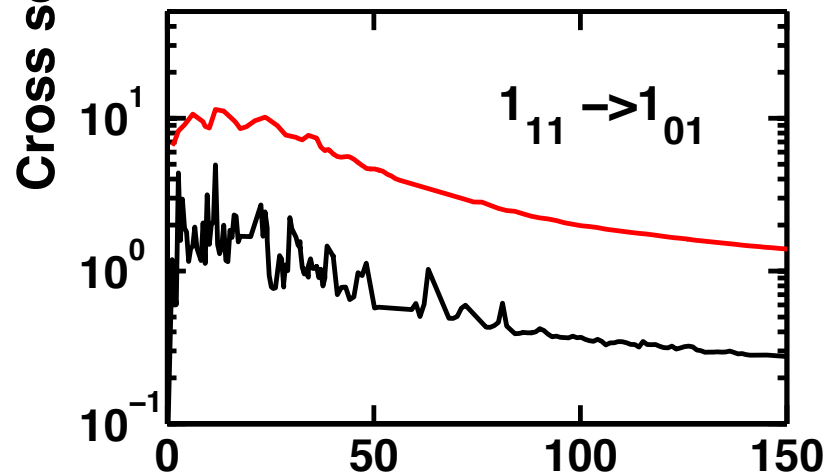
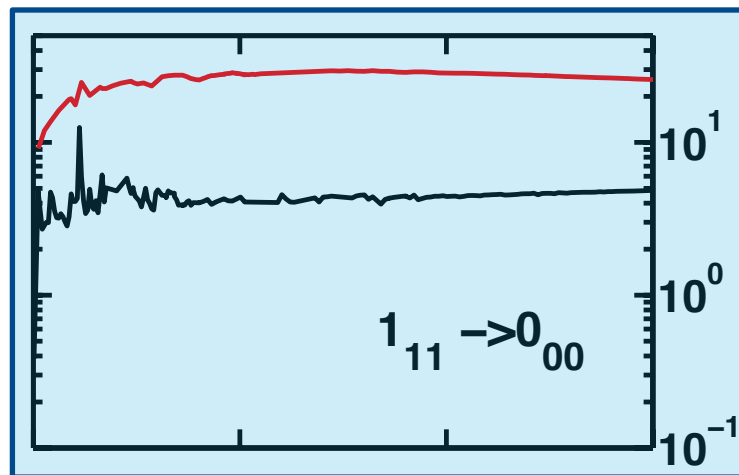
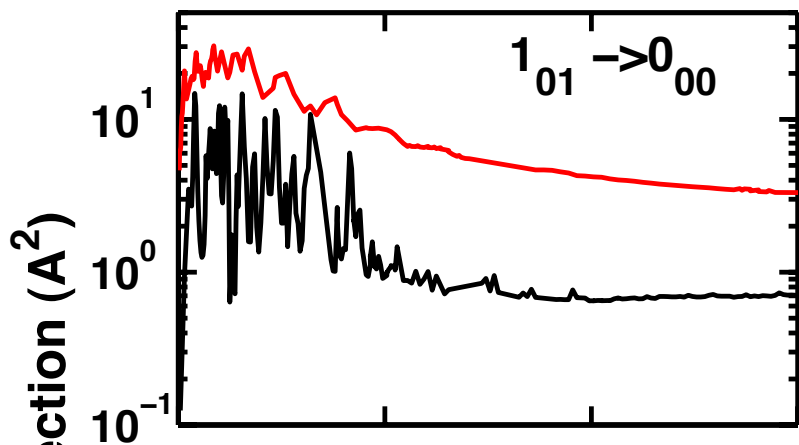
# Cross sections, HDO-H<sub>2</sub>

$$\sigma(E_{\text{collision}})$$

Ortho H<sub>2</sub>  $j=1$

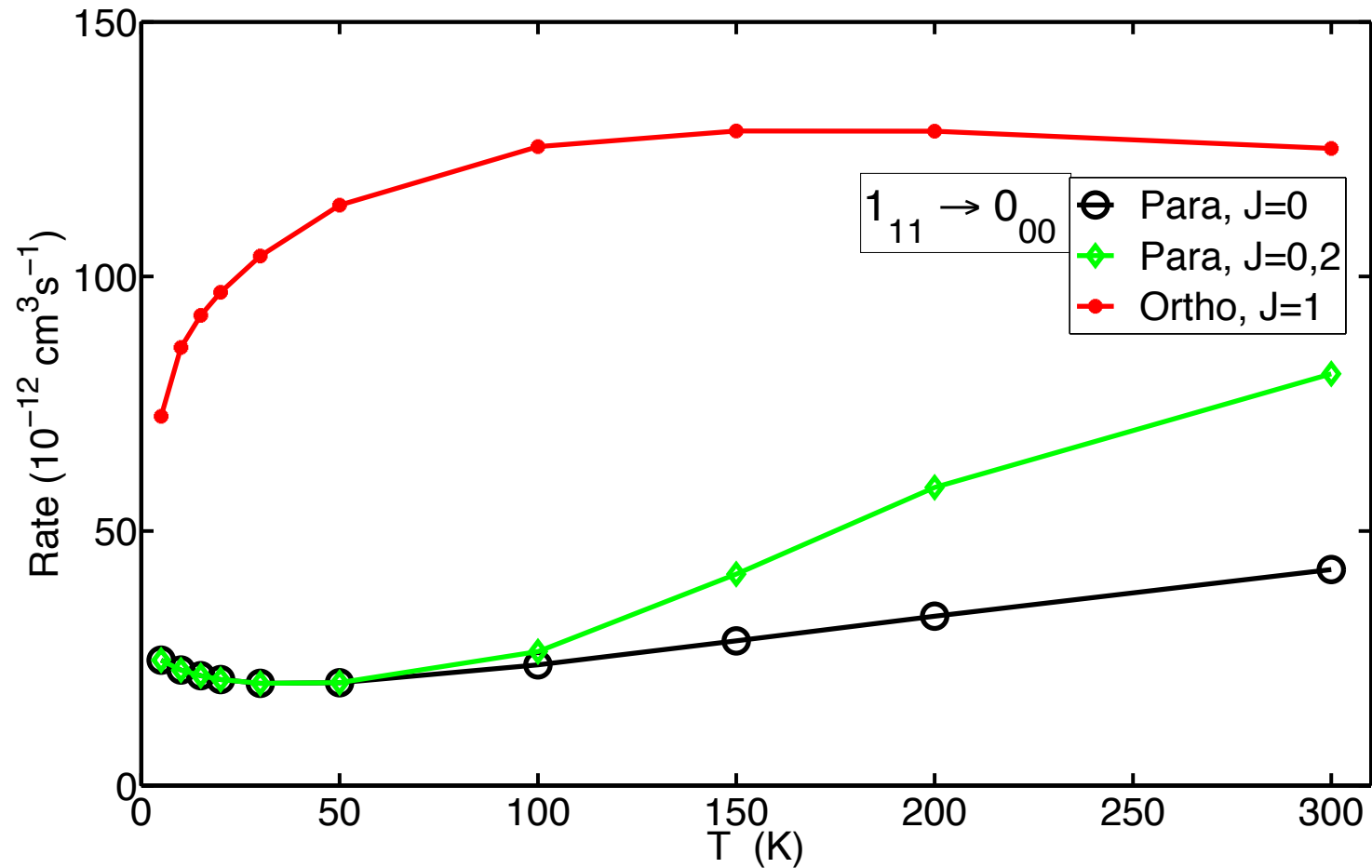


Para H<sub>2</sub>  $j=0,2$



Collision energy (cm<sup>-1</sup>)

# One example of rates



# The file... as in LAMDA :

## To be used in radiation transfer codes.

- HDO rates; ortho H<sub>2</sub>, J=1

- | i | f | 4 | 10     | 20     | 30     | 40     | 50     | 60     | 70     | 80     | 90     | 100    |        |
|---|---|---|--------|--------|--------|--------|--------|--------|--------|--------|--------|--------|--------|
| • | 2 | 1 | 68.70  | 65.50  | 55.87  | 48.09  | 42.38  | 38.14  | 34.93  | 32.41  | 30.40  | 28.76  | 27.40  |
| • | 3 | 1 | 61.94  | 83.57  | 94.97  | 101.74 | 107.44 | 112.19 | 115.99 | 118.95 | 121.23 | 122.96 | 124.26 |
| • | 3 | 2 | 64.48  | 78.22  | 75.16  | 67.86  | 61.19  | 55.68  | 51.18  | 47.46  | 44.37  | 41.75  | 39.51  |
| • | 4 | 1 | 7.28   | 10.98  | 10.77  | 9.47   | 8.23   | 7.20   | 6.36   | 5.67   | 5.11   | 4.63   | 4.23   |
| • | 4 | 2 | 112.91 | 180.29 | 212.04 | 223.60 | 229.98 | 233.84 | 236.04 | 237.09 | 237.33 | 236.99 | 236.26 |
| • | 4 | 3 | 51.15  | 71.42  | 71.53  | 67.06  | 62.82  | 59.21  | 56.15  | 53.52  | 51.24  | 49.25  | 47.49  |
| • | 5 | 1 | 17.30  | 20.70  | 21.41  | 21.50  | 21.63  | 21.78  | 21.91  | 22.02  | 22.12  | 22.21  | 22.29  |
| • | 5 | 2 | 35.26  | 42.49  | 42.03  | 40.16  | 38.60  | 37.31  | 36.22  | 35.29  | 34.48  | 33.78  | 33.17  |
| • | 5 | 3 | 102.42 | 125.31 | 129.32 | 127.62 | 125.08 | 122.13 | 119.04 | 115.95 | 112.96 | 110.13 | 107.46 |
| • | 5 | 4 | 49.04  | 51.33  | 47.62  | 44.29  | 41.67  | 39.49  | 37.60  | 35.95  | 34.49  | 33.19  | 32.04  |
| • | 6 | 1 | 5.98   | 7.58   | 8.37   | 8.97   | 9.37   | 9.63   | 9.80   | 9.91   | 9.99   | 10.04  | 10.07  |
| • | 6 | 2 | 71.93  | 86.73  | 93.27  | 100.90 | 107.12 | 111.70 | 115.00 | 117.39 | 119.12 | 120.40 | 121.35 |
| • | 6 | 3 | 35.88  | 41.20  | 39.63  | 38.56  | 37.62  | 36.68  | 35.77  | 34.90  | 34.09  | 33.36  | 32.69  |
| • | 6 | 4 | 53.64  | 64.95  | 67.11  | 68.04  | 67.74  | 66.64  | 65.16  | 63.53  | 61.89  | 60.29  | 58.76  |
| • | 6 | 5 | 57.87  | 61.02  | 52.66  | 46.89  | 42.45  | 38.77  | 35.67  | 33.02  | 30.76  | 28.80  | 27.10  |
| • | 7 | 1 | 7.32   | 8.34   | 9.45   | 10.40  | 10.93  | 11.19  | 11.31  | 11.34  | 11.34  | 11.30  | 11.26  |
| • | 7 | 2 | 12.09  | 13.70  | 13.38  | 12.90  | 12.28  | 11.64  | 11.06  | 10.54  | 10.09  | 9.69   | 9.34   |
| • | 7 | 3 | 57.84  | 64.92  | 74.44  | 83.60  | 89.32  | 92.71  | 94.70  | 95.82  | 96.42  | 96.67  | 96.72  |
| • | 7 | 4 | 18.71  | 23.10  | 24.24  | 24.83  | 24.91  | 24.73  | 24.43  | 24.10  | 23.76  | 23.43  | 23.11  |
| • | 7 | 5 | 121.05 | 139.28 | 158.27 | 176.12 | 187.14 | 193.54 | 197.13 | 199.01 | 199.82 | 199.93 | 199.59 |
| • | 7 | 6 | 49.41  | 57.06  | 57.04  | 55.79  | 53.45  | 50.80  | 48.23  | 45.86  | 43.74  | 41.85  | 40.17  |
| • | 8 | 1 | 0.13   | 1.18   | 2.34   | 2.90   | 3.22   | 3.43   | 3.58   | 3.69   | 3.79   | 3.87   | ....   |

A Faure et al  
MNRAS, 2012

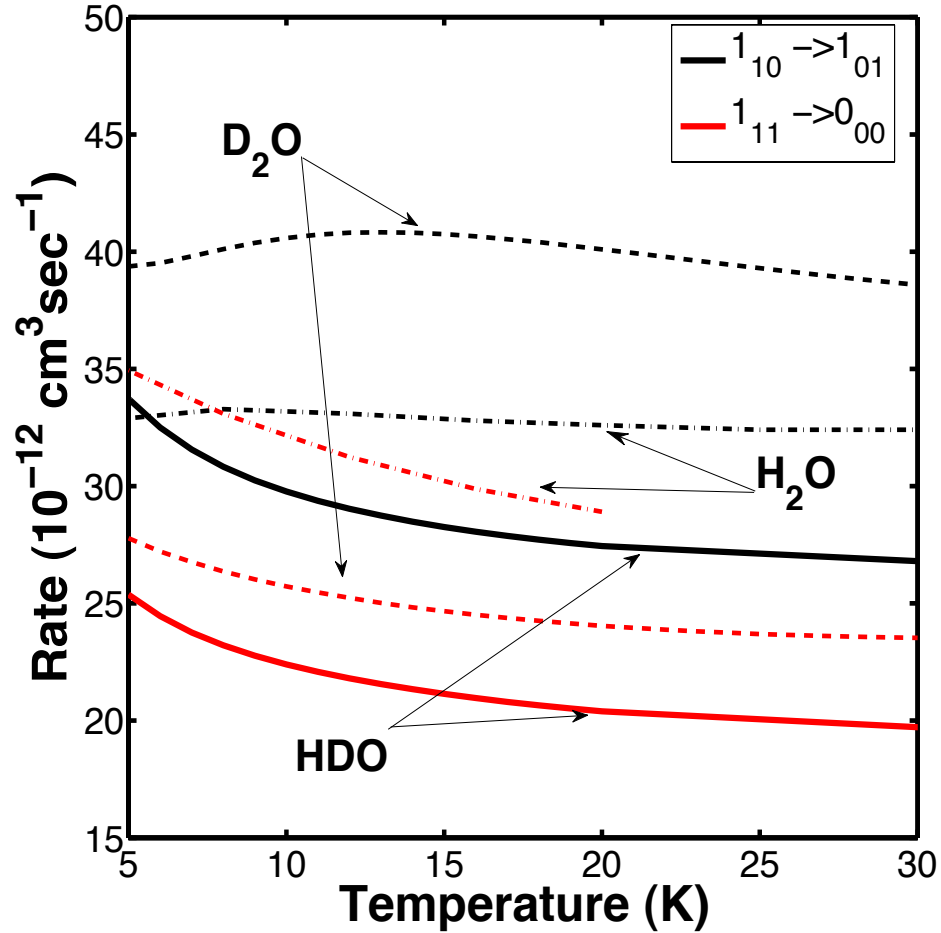
# Rates

Compare *red to red*, *black to black*..

## H<sub>2</sub>O vs HDO

HDO –H<sub>2</sub> rates similar to  
H<sub>2</sub>O – H<sub>2</sub> rates, but not  
quite...

*300K, o & p H<sub>2</sub>*  
*Effort 10-15 times as large  
as H<sub>2</sub>O*



# Using, testing the PES

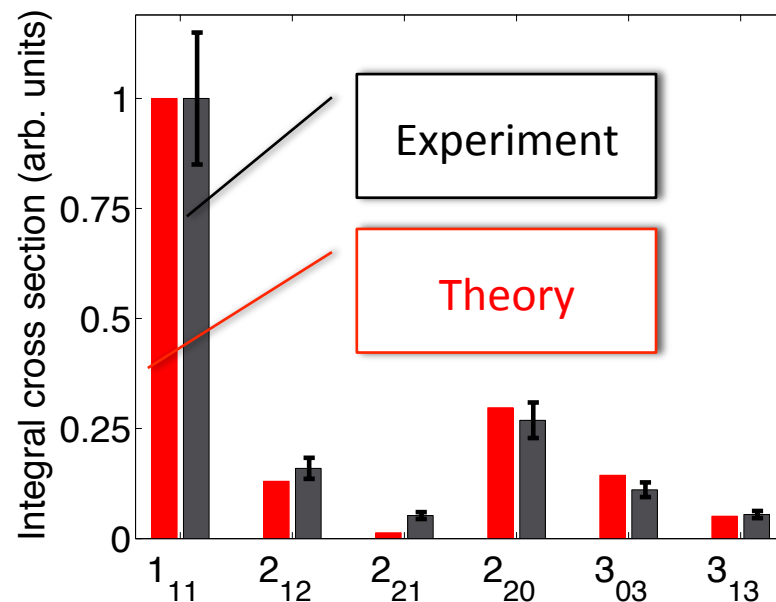
Property	Test	Status	
		Experiment	Theory
Overall shape	Pressure broadening	✓	✓
	Virial coefficients	?	☐
Wells	Bound states spectroscopy	✓	✓
Repulsive wells	Inelastic DCS	✓	✓
Vibrational transitions	Total cross section	1	In progress

# Experimental comparisons

- The same PES for each  $\text{H}_2\text{O}$ ,  $\text{HDO}$ , and  $\text{D}_2\text{O}$ . Test for one is valid for the others, except for

- Experimental procedures
- Specific numerical difficulties

- Tests:
  - Differential Cross Section
  - Pressure broadening
  - Integral Cross Section
  - Comparison ICS HDO – ortho  $\text{H}_2$



Sarma et al, JCP, 2013

Parker & LW, PCCP feature article, 2013, in prep.

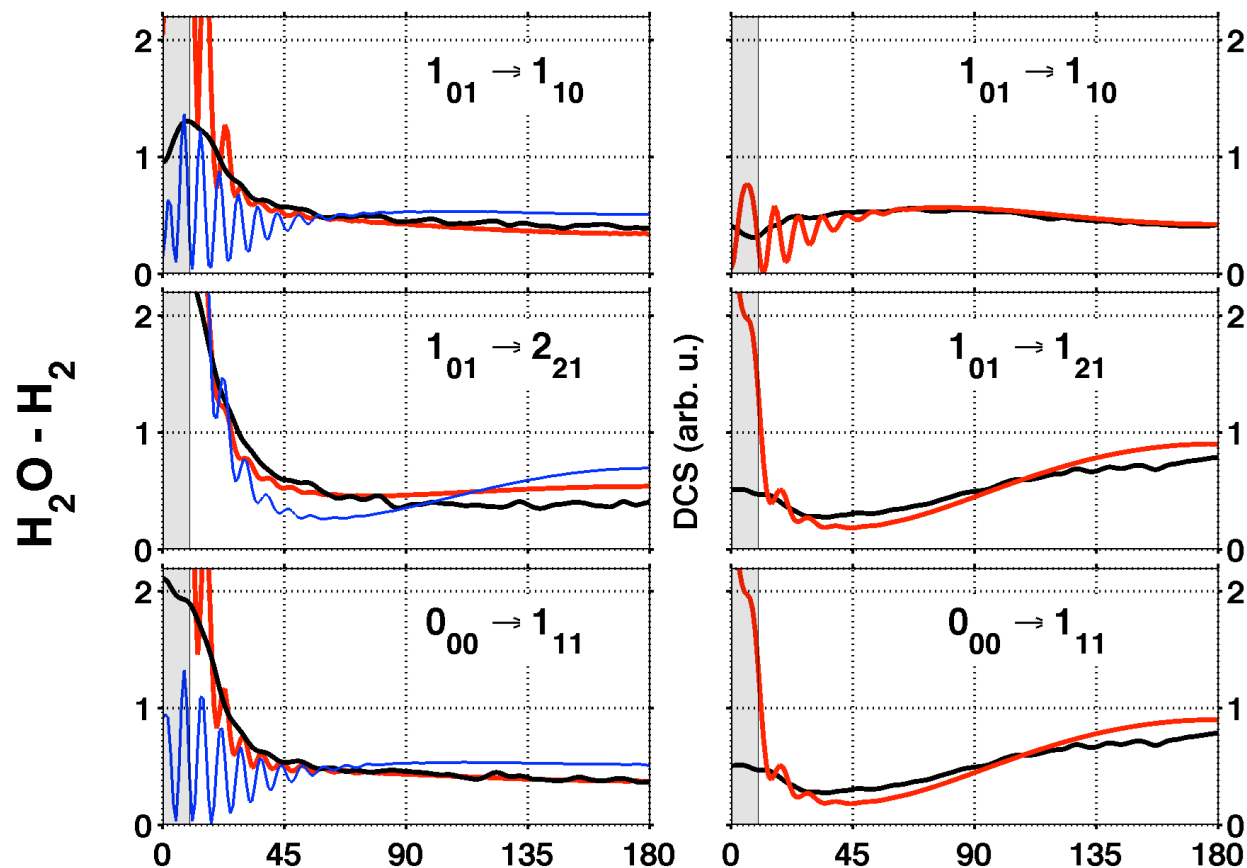
# H<sub>2</sub>O- H<sub>2</sub> DCS

Theory: 

Experiment: 

Normalized,  
no adjustable  
parameter  
(H<sub>2</sub> J=0,1,2, exp.  
abundances)

E collision : 574  
cm-1

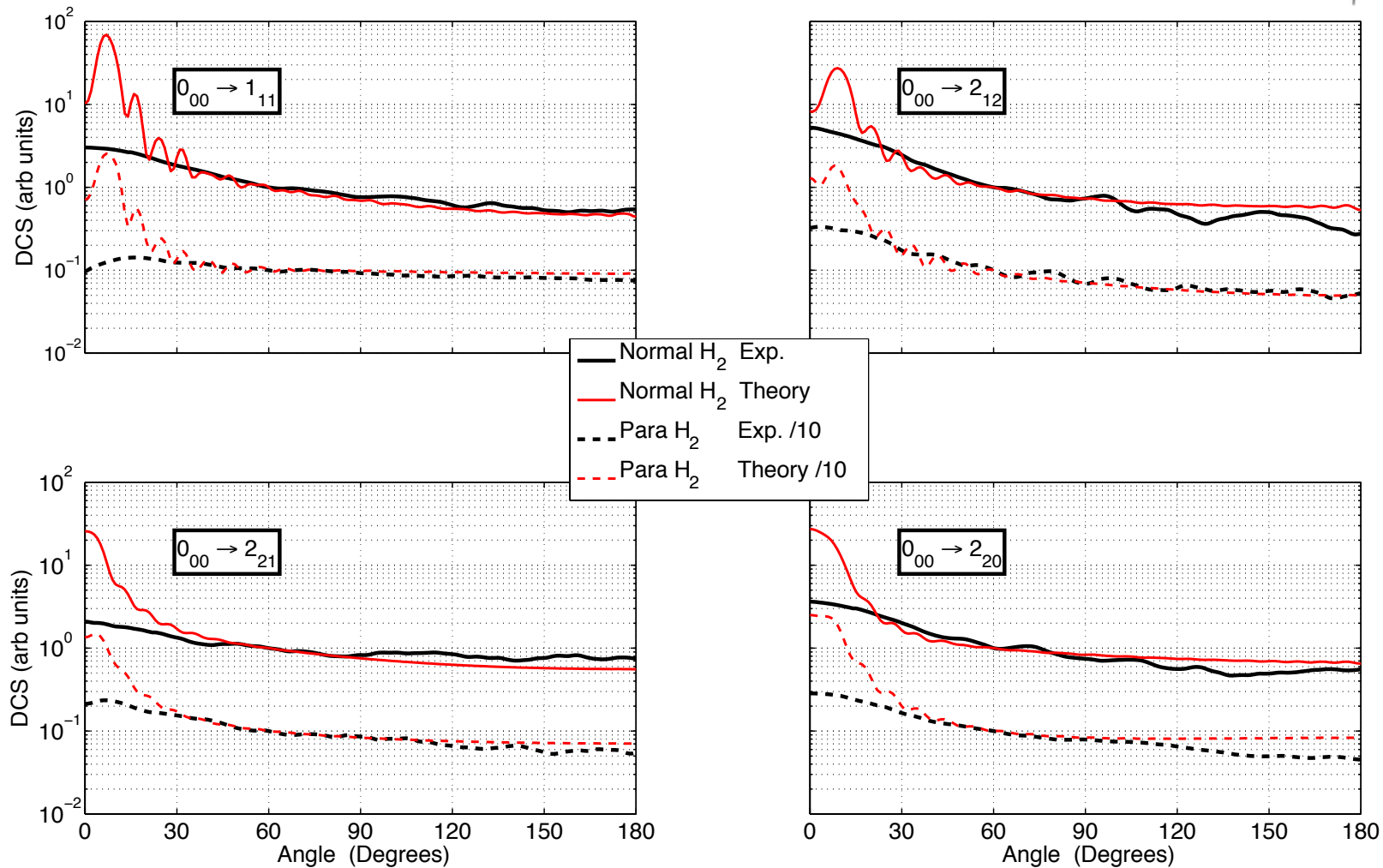


H<sub>2</sub>O - He

Garching MPE January 2013

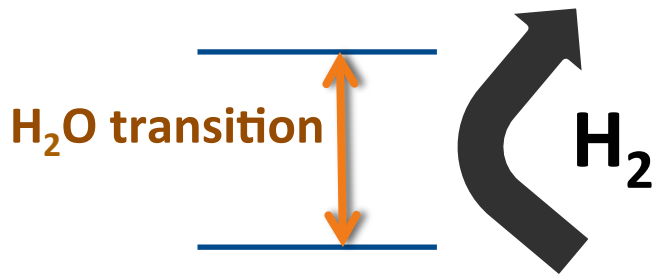


# HDO- H<sub>2</sub> DCS



# Pressure broadening

Theory



The collisions have two effects :

Virtual inelastic scattering  
Dephasing of the transition

Result in a spectral line profile +  
spectral line shift

Approximation of isolated collisions  
+ isolated spectallines.

## Experimental set up, Drouin et al, JPL

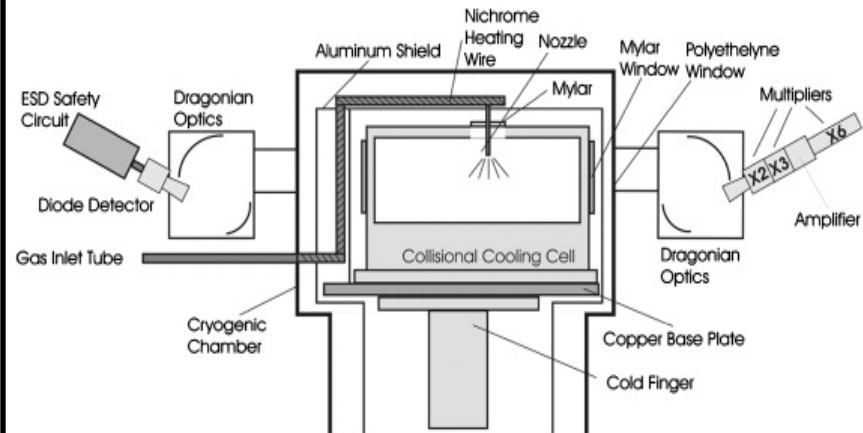
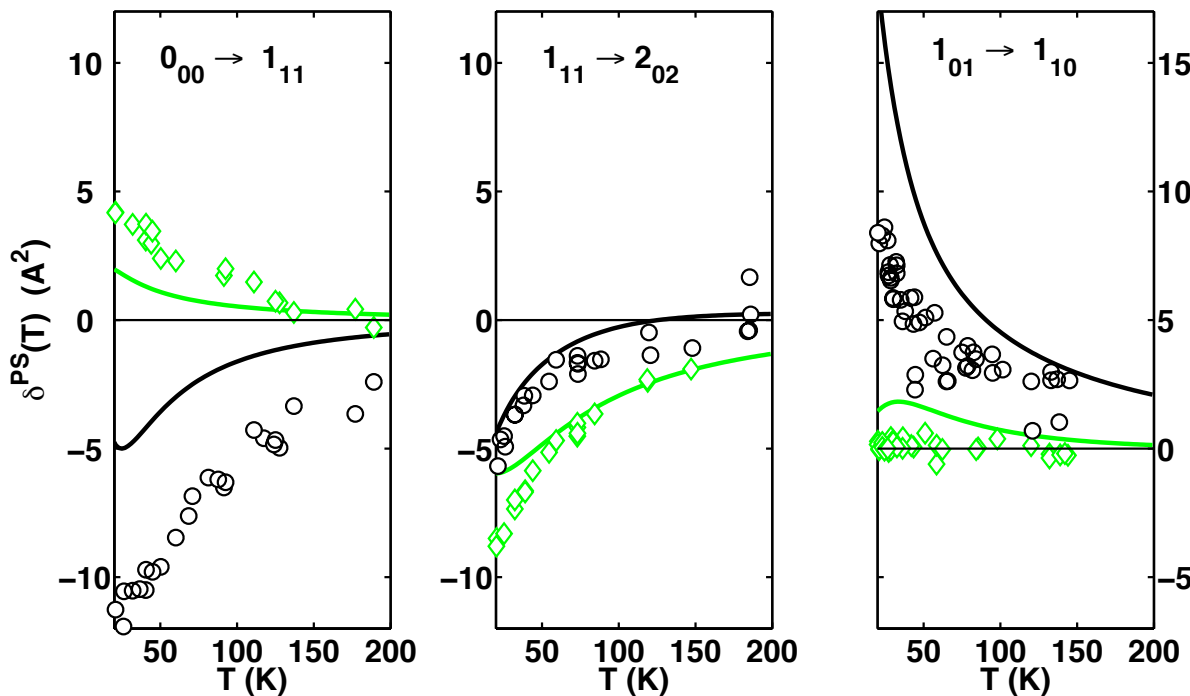
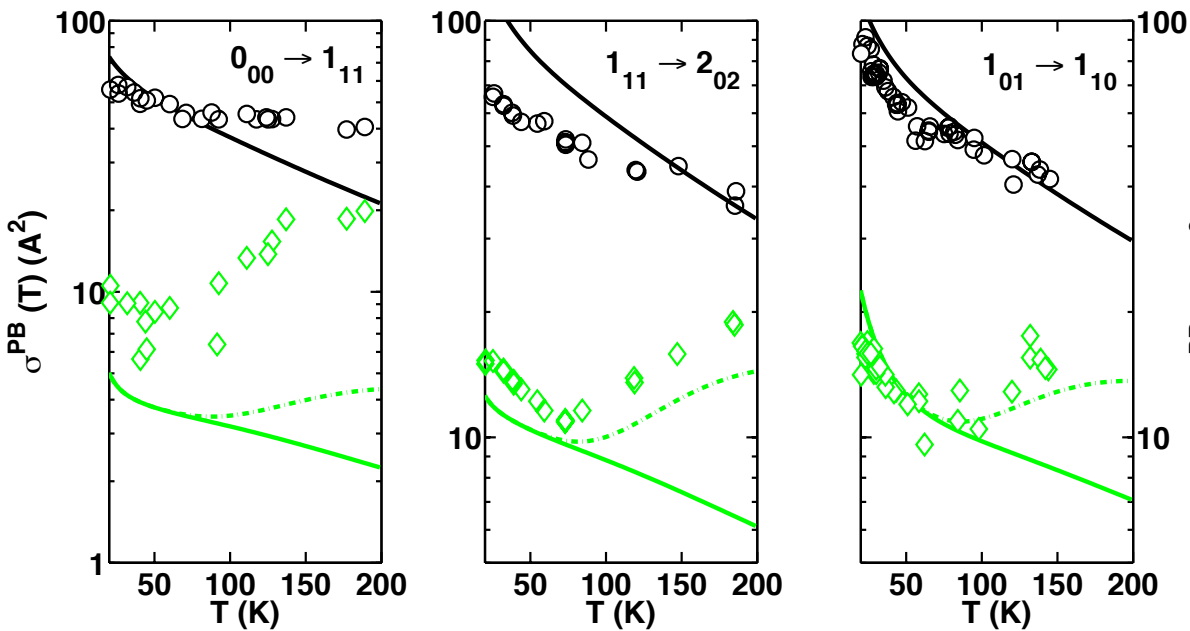


Fig. 1 Side-view of the collisional cooling experiment. The collisional cooling cell is housed inside the cryogen cell, cooled by a cold finger coupled to a helium cryo-refrigeration system and isolated from the warm cryogen cell walls via an Alu...

Michael J. Dick , Brian J. Drouin , Timothy J. Crawford , John C. Pearson



## Pressure broadening and shift, H<sub>2</sub>O in H<sub>2</sub>

B. Drouin (JPL), LW, Phys Rev A 2012.

BLACK: Broadening with normal H<sub>2</sub>;

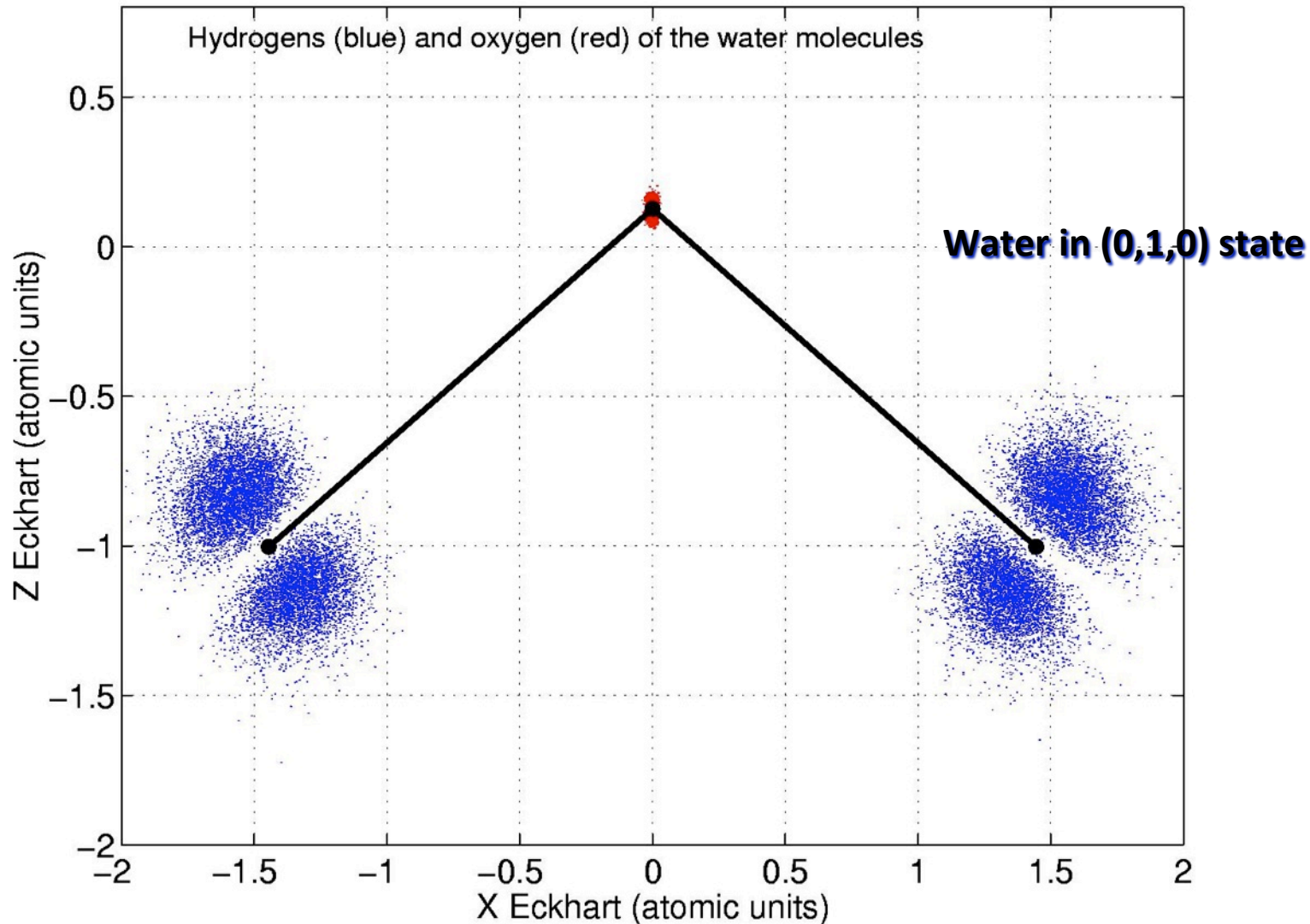
Green, broadening by para H<sub>2</sub>.

Lines : theory; symbols: experiments.

# Vibrational quenching :

$\langle v=1 | V | v=0 \rangle$  matrix elements

10,000 water molecules in the  $v=1$  state



# Vibrational quenching $v=1 \rightarrow v=0$

## H<sub>2</sub>O – o-H<sub>2</sub>

